

GENERAL CHEMISTRY

MODULE	CONTENT	YEAR	TERM	CREDITS	TYPE
I - Fundamentals	Chemistry	1	1	6	FB
LECTURER(S)			Postal address, telephone nº, e-mail address		
<ul style="list-style-type: none"> Juan José Díaz Mochón: Part I (Groups C y E) Professor non-designated: Part II (Groups C y E) 			Prof. Juan J. Díaz Mochón Dpto. Química Farmacéutica y Orgánica, Third floor, Faculty of Pharmacy. Pfizer - Universidad de Granada - Junta de Andalucía Centre for Genomics and Oncological Research (GENYO). Avenida de la Ilustración 114, Granada. Tel: +34 958 715 500, Ext 162 e-mail: juandiaz@ugr.es Professor non-designated Dpto. Físicoquímica, Second floor, Faculty of Pharmacy. e-mail:		
DEGREE WITHIN WHICH THE SUBJECT IS TAUGHT			TUTORSHIPS		
Human Nutrition and Dietetics			JJDM: Tuesday and Thursday from 16:00 h to 18:00 h; Friday from 13:30 h to 15:30 h PND:		
PREREQUISITES and/or RECOMMENDATIONS (if necessary)					
<ul style="list-style-type: none"> Suitable knowledge of Chemistry, high-school level. Basic knowledge on mathematical operations (logarithmic and exponential functions, use of calculator, etc.). 					
BRIEF ACCOUNT OF THE SUBJECT PROGRAMME (ACCORDING TO THE DEGREE ¿??)					
GENERAL AND PARTICULAR ABILITIES					
General: CG3, CG4, CG8, CG29 Specific: CE1, CE10, CE11, CE13.					



OBJECTIVES (EXPRESSED IN TERMS OF EXPECTED RESULTS OF THE TEACHING PROGRAMME)

DETAILED SUBJECT SYLLABUS

LECTURES:

PART I

- UNIT 1. **Introduction to Organic Chemistry:** Concept and evolution. Carbon-carbon and carbon-heteroatom bonds. Intermolecular forces. Hydrocarbons: classification and nomenclature. Constitutional isomerism. Stereoisomerism.
- UNIT 2. **Functional Groups.** Classification of Organic Compounds. Nomenclature of functional groups. Concept of functional groups. Principal monovalent functional groups: alcohols, phenols, ethers and amines. Principal divalent functional groups: aldehydes, ketones and imines. Principal trivalent functional groups: acids, esters, amides and nitriles.
- UNIT 3. **Carbohydrates.** Overview of Carbohydrates. Classification and nomenclature. Acyclic and cyclic forms. Type of Representation: Fischer and Haworth. Reactivity of monosaccharides: oxidation, reduction, O- and N-glycosylation. Modification of monosaccharides: amino sugars and deoxy sugars. Classification of disaccharides. Main disaccharides. Oligo and polysaccharides.
- UNIT 4. **Lipids.** Overview of Lipids. Oils, fats and waxes. Phospholipids and sphingolipids. Structure of membrane lipids. Eicosanoids: arachidonic acid. Isoprene and Isoprenoids. Steroids.
- UNIT 5. **Amino acids, peptides and proteins.** Amino acids: Structure and classification. Properties synthesis and protection of amino acids. Peptides and proteins. Primary and secondary structures of proteins. Enzymes and cofactors.
- UNIT 6. **Vitamins.** Overview of vitamins. Hidrosoluble and liposoluble vitamins. Biological functions.

PART II

- UNIT 7. **Structure, chemical bonds, and reactivity.** Basic principles. Ionic networks. Covalent bond. Metallic bond. Intermolecular forces and biological role. Chemical reactions. Limiting reagent. Reaction yield.
- UNIT 8. **Solutions.** Classification. Concentration. Intermolecular forces and solubility. Colligative properties and applications.
- UNIT 9. **Thermochemistry.** First law of Thermodynamics. Heat. Work. Internal energy. Enthalpy. Heat capacity. Enthalpy of physical changes. Reaction enthalpy. Determination of reaction enthalpy. Hess' law. Influence of the temperature on the reaction enthalpy. Food and energy resources.
- UNIT 10. **Equilibrium.** Spontaneous processes and equilibrium. Second law of Thermodynamics. Gibbs' free energy. Phase equilibrium and phase diagrams. Chemical equilibrium and law of mass action. Thermodynamic origin of the equilibrium constant. Le Chatelier's principle. Effect of pressure, concentration, and temperature on the equilibrium.
- UNIT 11. **Proton transfer reactions. Acid-base equilibria.** Strength of acids and bases. Dissociation constants. Water autoprotolysis. pH scale. pH determination in simple systems. Acid-base titrations. Buffers. Biological role of the acid-base equilibria.
- UNIT 12. **Electron transfer reactions. Redox equilibria.** Oxidation and reduction. Balancing redox reactions. Standard reduction potentials and equilibrium constant. Spontaneity and reaction direction. Applications and redox systems of biological importance. Antioxidants.

LABORATORY SESSIONS AND SEMINARS:

Seminars

- Problems solving

Laboratory sessions

- **Session 1. Introduction to laboratory. Solutions. Preparation of a buffer solution.**
- **Session 2. Acidity measurements of olive oil, milk, and vinegar.**
- **Session 3. Practical test: Acid – base titration.**



- **Session 4. *Synthesis of Isoamyl acetate.***
- **Session 5. *Synthesis of dibenzalacetone.***
- **Session 6. *Hydrolysis of sucrose.***

READING

BASIC READINGS:

- **Chemical Principles.** 3^a-5^a Eds. P. Atkins, L. Jones.
- **General chemistry.** 10^a Ed. R.H Petrucci, F. G. Herring, J. D. Madura, C. Bissonnette.
- **Química Orgánica.** H. Hart, L. E. Crine, D. J. Hart y Ch. M. Hadad. Ed. McGraw Hill. 12^a edición., 2007.
- **Química de los alimentos.** H. D. Belitz. Ed. Acribia, 2011.

COMPLEMENTARY READINGS:

- **General chemistry.** J. L. Rosenberg, L. M. Epstein.
- **Resolución de Problemas de Química.** A. Sánchez Coronilla.
- **Resolución de Problemas de Química General.** C. J. Willis.
- **Química. Un proyecto de la American Chemical Society.** Varios autores. Ed. Reverté. 2005.
- **Cuestiones y ejercicios de Química Orgánica. Una guía de autoevaluación.** E. Quiñoá y R. Riguera (2^a Ed.) Ed. Mc Graw Hill 2004.
- **Nomenclatura y representación de los compuestos orgánicos (Una guía de estudio y autoevaluación).** E. Quiñoá y R. Riguera. Ed. Mc Graw Hill 2005.

